



9) **Explain** each answer briefly.

a) Place the following elements in order of increasing ionization energy: F, O, and S.

b) Which has the largest ionization energy: O, S, or Se?

c) Which has the greatest electron affinity: Se, Cl, or Br?

d) Which has the largest radius:  $O^{2-}$ , F, or F

e) Rank the following in order of increasing atomic radius: O, S, and F.

f) Which has the largest ionization energy: P, Si, S, or Se?

g) Place the following in order of increasing radius: Ne,  $O^{2-}$ ,  $N^{3-}$ , or F.